

	SURFACE VEHICLE INFORMATION REPORT	
	SAE	J1498 DEC2011
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Superseding J1498 DEC2005		
Heating Value of Fuels		

RATIONALE

The heating value of fuel is a measure of its heat of combustion or energy content as determined under specific conditions. Higher (gross) and lower (net) heating values are discussed. Gross (liquid water) and net (gaseous water) heating values differ by the heat of vaporization of the water in the combustion product mixture. This SAE Recommended Practice discusses the measurement and calculated estimated heating values for gaseous, liquid, and solid fuels.

1. SCOPE

The heating value or heat of combustion is a measure of the energy available from the fuel. The fraction or percentage of the heat of combustion that is converted to useful work is a measure of the thermal efficiency of an engine. Thus, a knowledge of the heat of combustion of the fuel is basic to the engineering of automotive engines. This SAE Information Report provides information on the standardized procedures for determining the heat of combustion of fuels that may be used for automotive engines.

The changes to SAE J1498 include:

SAE Publications - Added SAE Paper 2010-01-1517

Other Publications and Sections 5, 9, and 10 - Updated ASTM alphanumeric designations and titles.

Section 10 - Added discussion of a method to calculate net heating value for gasoline-ethanol blends using ASTM D3338.

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http://www.sae.org/technical/standards/J1498_201112**

2. REFERENCES

2.1 Applicable Documents

The following publications form a part of this specification to the extent specified herein. Unless otherwise indicated, the latest issue of SAE publications shall apply.

2.1.1 SAE Publications

Available from SAE International, 400 Commonwealth Drive, Warrendale, PA 15096-0001, Tel: 877-606-7323 (inside USA and Canada) or 724-776-4970 (outside USA), www.sae.org.

1. Pat Geng, Robert L. Furey, and i.V. Anke Konzack, "Calculation of Heating Value for Gasoline Containing Ethanol." SAE Paper 2010-01-1517.

SAE J313 Diesel Fuels

SAE TSB 003 Rules for SAE Use of SI (Metric) Units

2.2 Other Publications

1. J. M. Sturtevant, "Calorimetry", Chapter VII, Vol. I, Part V of "Physical Methods of Chemistry", Edited by A. Weissberger and B. W. Rossiter, Wiley-Interscience, New York, NY, 1971.
2. ASTM Test Methods and Practices cited throughout the text are available in the latest volume of the Annual Book of ASTM Standards as shown as follows. These volumes are available from ASTM International, 100 Barr Harbor Drive, P.O. Box C700, West Conshohocken, PA 19428-2959, Tel: 610-832-9585, www.astm.org

ASTM	Volume	Title
D86	05.01	Test Method for Distillation of Petroleum Products at Atmospheric Pressure
D240	05.01	Test Method for Heat of Combustion of Liquid Hydrocarbon Fuels by Bomb Calorimeter
D1018	05.01	Test Method for Hydrogen In Petroleum Fractions
D1405	05.01	Test Method for Estimation of Net Heat of Combustion of Aviation Fuels
D1826	05.06	Test Method for Calorific (Heating) Value of Gases in Natural Gas Range by Continuous Recording Calorimeter
D1945	05.06	Test Method for Analysis of Natural Gas by Gas Chromatography
D1946	05.06	Practice for Analysis of Reformed Gas by Gas Chromatography
D2421	05.01	Practice for Interconversion of Analysis of C5 and Lighter Hydrocarbons to Gas-Volume, Liquid-Volume, or Mass Basis
D2650	05.01	Method for Chemical Composition of Gases By Mass Spectrometry
D3338	05.01	Test Method for Estimation of Net Heat of Combustion of Aviation Fuels
D3588	05.01	Practice for Calculating Heat Value, Compressibility Factor, and Relative Density of Gaseous Fuels
D3701	05.01	Test Method for Hydrogen Content of Aviation Turbine Fuels by Low Resolution Nuclear Magnetic Resonance Spectrometry
D4529	05.02	Test Method for Estimation of Net Heat of Combustion of Aviation Fuels
D4809	05.02	Test Method for Heat of Combustion of Liquid Hydrocarbon Fuels by Bomb Calorimeter (Precision Method)
D4868	05.02	Test Method for Estimation of Net and Gross Heat of Combustion of Burner and Diesel Fuels
D5865	05.02	Test Method for Gross Calorific Value of Coal and Coke
SI 10	14.04	American National Standard for Metric Practice
E711	11.04	Test Method for Gross Calorific Value of Refuse-Derived Fuel by the Bomb Calorimeter

3. F.D. Rossini, K. S. Pitzer, R. L. Arnett, R. M. Braun, and G. C. Pimental, "Selected Values of Physical and Thermodynamic Properties of Hydrocarbons and Related Compounds", pp. 445-463, Carnegie Press, Carnegie Institute of Technology, Pittsburgh, PA, 1953.
4. ASTM Data Series Publication DS4B, "Physical Constants of Hydrocarbon and Non-Hydrocarbon Compounds", ASTM International, 100 Barr Harbor Drive, P.O. Box C700, West Conshohocken, PA 19428-2959, Tel: 610-832-9585, www.astm.org, 1991.
5. "Reference Data for Hydrocarbons and Petro-Sulfur Compounds", Phillips Petroleum Company Bulletin No. 521, Bartlesville, Oklahoma.
6. R. C. Wilhoit and B. J. Zwolinski, "Physical and Thermodynamic Properties of Aliphatic Alcohols", Journal of Physical and Chemical Reference Data, Vol. 2, Supplement No. 1, 1973, American Chemical Society and the American Institute of Physics.
7. E. S. Domalski, "Selected Values of Heats of Combustion and Heats of Formation of Organic Compounds Containing the Elements of C, H, N, O, P, and S", Journal of Physical and Chemical Reference Data, Vol. 1, No. 2, pp. 221-278, 1972, American Chemical Society and the American Institute of Physics.
8. "Technical Data Book - Petroleum Refinery", 3rd Edition, American Petroleum Institute, Washington, DC, 1977.
9. J. D. Cox and G. Pilcher, "Thermochemistry of Organic and Organometallic Compounds", Academic Press, 1970.

3. MEASUREMENT OF HEAT OF COMBUSTION

Heats of combustion can be determined to high precision by burning a known amount of the sample completely in pure oxygen in a calibrated calorimeter. The heat of combustion is the product of the temperature rise and the total heat capacity of the calorimeter and its contents. For precision measurements, corrections are applied for energy added to initiate combustion and for the energy involved in the formation of extraneous products such as nitric acid and sulfuric acid.

Calorimeters of various types have been developed to obtain heats of combustion of gases, liquids, or solids to various levels of precision (Reference 1). For evaluation of fundamental thermodynamic properties, precision of calorimetric measurements of the highest order is required since the results are obtained as small differences in large quantities. For heats of combustion of fuel mixtures, lower precisions are usually sufficient for many purposes and less sophisticated techniques are usually used.

Heats of combustion or calorific values of gaseous fuels are usually measured in flame calorimeters in which the sample and oxygen react at constant pressure. Heats of combustion of solid and liquid fuels are usually measured in bomb calorimeters at constant volume.

4. UNITS OF MEASUREMENT

In the SI system of measurement, the heat of combustion is expressed in joules per kilogram. For materials generally used as fuels, the heat of combustion is of such magnitude that it is customarily expressed in terms of a multiple of the basic unit, namely in megajoules per kilogram, MJ/kg. Where the molecular weight is known as in pure compounds, it is sometimes expressed in terms of megajoules per mole, MJ/mole.

A large body of literature is available - and still accumulating - in terms of other, non-SI units such as calories per gram or British Thermal Units (Btu) per pound. The following factors can be used in converting these other units to SI units as in Table 1:

TABLE 1 - UNITS OF MEASUREMENT

To Convert From	To	Multiply By
Calorie (International Table)	Joule (J)	4.186 800
British Thermal Unit (IT)	Joule (J)	1055.056
Calorie/Gram	Megajoules/kilogram (MJ/kg)	0.004 186 800
Btu/pound	Megajoules/kilogram (MJ/kg)	0.002 326 000
Pound (Avoirdupois)	Kilogram (kg)	0.453 592 4
Btu/cubic foot	Megajoules/cubic meter (MJ/m ³)	0.037 258 95

SAE TSB 003 and ASTM SI 10 list conversion factors for units other than those included here. In any conversions, the number of digits retained should not exceed that implied by the precision of the original quantity.

Many fuels are sold by volume. Heating values are, therefore, frequently expressed in terms of heat units on a volume basis. Thus, the heating values of gaseous fuels are given in terms of heat units per cubic foot or per cubic meter and those of liquid fuels are frequently given in terms of heat units per gallon or per liter. These terms can be computed to a mass basis using the density in appropriate units at the appropriate temperature and pressure.

5. TYPES OF HEATING VALUES

Four types of heating values may be defined according to the assumed state of water as a combustion product (gaseous or liquid) and according to the nature of the assumed combustion process (constant volume or constant pressure). Gross (liquid water) and net (gaseous water) heating values differ by the heat of vaporization of the water in the combustion product mixture. Constant volume and constant pressure heating values differ by the heat equivalent of the work represented by the volume change at constant pressure caused by the change in the number of moles during combustion. Numerical differences between constant volume and constant pressure combustion are small and normally neglected for fuels. Tables of heating values usually list gross and net values at constant pressure.

ASTM D4809 defines the gross heat of combustion, Q_g , as follows:

"The gross heat of combustion at constant volume of a liquid or solid fuel containing only the elements carbon, hydrogen, oxygen, nitrogen, and sulfur is the quantity of heat liberated when a unit mass of the fuel is burned in oxygen in an enclosure of constant volume, the products of combustion being gaseous carbon dioxide, nitrogen, sulfur dioxide, and liquid water, with the initial temperature of the fuel and the oxygen and the final temperature of the products at 25 °C."

ASTM D4809 also defines the net heat of combustion, Q_n , as:

"The net heat of combustion at constant pressure of a liquid or a solid fuel containing only the elements carbon, hydrogen, oxygen, nitrogen, and sulfur is the quantity of heat liberated when a unit mass of the fuel is burned in oxygen at a constant pressure of 0.101 MPa (1 atm), the products of combustion being carbon dioxide, nitrogen, sulfur dioxide, and water, all in the gaseous state, with the initial temperature of the fuel and the oxygen and the final temperature of the products of combustion at 25 °C."

These two types are interrelated by Equation 1:

$$Q_n (\text{net, } 25^\circ\text{C}) = Q_g (\text{gross, } 25^\circ\text{C}) - 0.2122 \times H \quad (\text{Eq. 1})$$

where:

Q_n (net, 25°C) = net heat of combustion at constant pressure, MJ/kg

Q_g (gross, 25°C) = gross heat of combustion at constant volume, MJ/kg

H = mass % hydrogen in sample

NOTE: The percentage of hydrogen in the sample may be determined in accord with ASTM D1018 or with ASTM D3701.

The net heat of combustion is also known as the "lower heating value, (LVH)". It is the value customarily used in calculating thermal efficiency of engines since the exhaust gases are emitted at nearly the same pressure as that of the air and fuel supplied to the inlet system and since they contain water in the uncondensed or vapor state.

6. STANDARD CONDITIONS

Whenever heats of combustion are reported, the temperature, pressure, and physical state (solid, liquid, gas) of the reactants and products should be included. By international convention, a standard state of 25°C and 1 atm pressure is currently in use. It should be noted that in the change to the SI system of measurement, the standard state pressure is being changed to 1 bar, which is 100 kPa, instead of 1 atm which is 101.325 kPa. Few data are yet available at 1 bar pressure. This change will have negligible effect on the practical heating value of fuels.

Many data in the technical literature are given at temperatures other than the standard reference temperature of 25°C . These can be converted to 25°C (77°F) by correcting for the difference between the total heat capacity of the products of the combustion and the total heat capacity of the reactants at constant pressure. ASTM D5 4B, p. 62 (Reference 4) provides the details of such corrections when the heat capacities are known. ASTM D4809 provides a series of empirical correction factors that can be used for petroleum fuels when the heat capacities are unknown. Equation 2 applies to such corrections:

$$Q_g (\text{gross, } 25^\circ\text{C}) = Q_g (\text{gross, } t^\circ\text{C}) + A(t^\circ\text{C} - 25^\circ\text{C}) \quad (\text{Eq. 2})$$

where:

A = correction factor

Although heating values of gaseous fuels are usually expressed in terms of heating units per cubic foot, the amount of gas contained in a cubic foot has not been completely standardized throughout the gas industry. In ASTM D1826, the standard cubic foot is regarded to be the quantity of gas that will fill one cubic foot at a standard pressure of 14.73 psia (101.6 kPa) and a standard temperature of 15.56°C (60°F) when in equilibrium with water.

7. HEATING VALUES OF PURE COMPOUNDS

The heats of combustion of pure compounds, particularly hydrocarbons, have been measured with great precision because of their significance in elucidating structures and the chemical, physical, and thermodynamic properties of organic molecules. Major compilations of such data have been published (References 2 to 9) showing centimeter-gram-second metric and/or U.S. inch-pound units at standard state conditions of 25°C and 1 atm pressure.

8. HEATING VALUES OF MIXTURES

Heating values of mixtures of nonpolar liquids and gases, which do not react, are generally additive for practical purposes and can be computed from the known amounts and heating values of each component. When any of the components are polar liquids, solids, or gases that interact with each other, the heats of combustion may no longer be additive and, therefore, should be measured calorimetrically.

If water is present in the fuel as is frequently the case with such fuels as oxygenated fuels and coals, the total net heating value of the wet fuel will be lower than that of dry fuel. In this case, some of the heat of burning the fuel will be spent in evaporating the water present with the fuel.

9. STANDARDIZED METHODS FOR MEASURING HEATING VALUES

A variety of methods for measuring or estimating the heating values of fuels have been standardized by technical societies. The methods published by ASTM include statements of precision which are included here as guides to the usefulness of the method in proposed applications. A statement on precision, i.e., repeatability and reproducibility, provides information on the degree of mutual agreement between individual measurements while a statement on accuracy provides information on the agreement of the measurements with an accepted reference value. In general, 95% of the measurements can be expected to lie within the precision limits stated in the method when conducted properly. However, readers are cautioned to examine each of the methods for full information including the range of materials included in the interlaboratory comparison on which the method is based.

It should also be noted that the precision of the methods for computing heats of combustion from measured values of other properties such as aniline point and gravity are dependent upon the precision of these measured values rather than on the computation process.

9.1 Gaseous Fuels

The heat of combustion of a gaseous fuel can be measured directly or can be calculated from the composition of the mixture and the heats of combustion of its individual constituents provided that there is no chemical reaction between the constituents.

ASTM D1826 is a standardized procedure for measuring the heat of combustion of gases in the natural gas range of 900 to 1200 Btu/standard ft³ with a precision of 0.3 to 0.5%. In this procedure, the test gas is burned in a stream of air and the temperature rise of the air stream is measured. The water is condensed to the liquid state and the gross heat of combustion is obtained.

ASTM D3588 is applicable to all common types of utility gaseous fuels for which suitable methods of analysis are available. This includes dry natural gas, reformed gas, oil gas (both high and low Btu), propane-air, carbureted water gas, and coke oven and retort coal gas. The calorific values are calculated from the molar composition and the respective ideal gas values for the components. These values must then be converted to the real gas basis by means of compressibility summation factors which are listed in ASTM D3588 for most of the lower molecular weight hydrocarbons and gaseous nonhydrocarbons.

The following ASTM documents are helpful in determining the molar composition of the gaseous fuels:

- a. ASTM D1945 - Test Method for Analysis of Natural Gas by Gas Chromatography
- b. ASTM D1946 - Practice for Analysis of Reformed Gas by Gas Chromatography
- c. ASTM D2421 - Practice for Interconversion of Analysis of C₅ and Lighter Hydrocarbons to Gas-Volume, Liquid-Volume, or Weight Basis
- d. ASTM D2650 - Test Method for Chemical Composition of Gases by Mass Spectrometry

9.2 Liquid Fuels

Heats of combustion of liquid fuels and solid fuels as well, are usually measured in a bomb calorimeter equipped with either an isothermal or an adiabatic type jacket.

- a. ASTM D240 describes a method which is particularly applicable to hydrocarbons of both low and high volatility. The precision of this method is as follows:
 1. Repeatability - 0.13 MJ/kg (55 Btu/lb)
 2. Reproducibility - 0.40 MJ/kg (175 Btu/lb)
- b. ASTM D4809 was designed specifically for use with aviation turbine fuels but is also directly applicable to such fuels as gasolines, kerosenes, No. 1 and 2 fuel oil, No. 1-D and 2-D diesel fuel and No. 0-GT, 1-GT, and 2-GT gas turbine fuels. The precision of the method varies with the type of fuel and can be summarized in Table 2:

TABLE 2 - ASTM D4809 TEST METHOD PRECISION

	Gross Heat of Combustion MJ/kg	Gross Heat of Combustion Btu/lb	Net Heat of Combustion MJ/kg	Net Heat of Combustion Btu/lb
Repeatability:				
All Fuels	0.097	42	0.096	41
Nonvolatile Fuels	0.096	41	0.099	43
Volatile Fuels	0.100	43	0.091	39
Reproducibility:				
All Fuels	0.228	98	0.324	139
Nonvolatile Fuels	0.239	103	0.234	101
Volatile Fuels	0.207	89	0.450	193

9.3 Solid Fuels

ASTM has standardized three procedures for the determination of the gross heating value of solid fuels:

1. ASTM D5865 - Test Method for Gross Calorific Value of Coal and Coke
2. ASTM E711 - Test Method for Gross Calorific Value of Refuse-Derived Fuel by the Bomb Calorimeter

The precision of these three methods is listed in Table 3:

TABLE 3 - TEST METHOD PRECISION

	Repeatability MJ/kg	Repeatability Btu/lb	Reproducibility MJ/kg	Reproducibility Btu/lb
ASTM D5865				
Bituminous	0.160	69	0.249	107
Subbituminous-Lignite	0.140	59	0.326	140
ASTM E711	not determined	see standard	not determined	see standard